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 H O H O H O number of moles of unknown substance O CO H O C H CO H O C H O H O C H O CO use mole ratio 5 mol O 2 1 mol C 3 H 8 3 mol CO 2 1
 mol C 3 H 8 4 mol H 2 O 1 mol ...MATH SKILLS TRANSPARENCY MASTER 16 MATH SKILLS TRANSPARENCY WORKSHEET Determining Numbers of Ions
 Use with Chapter 7, Section 7.3 1. Write a simple word equation that illustrates what must be true of total positive charge and total negative charge in
 an ionic compound. 2. How many potassium ions (group 1) would be related to balance the charge of each of snow.dearbornschools.org T102
 Chemistry: Matter and Change Name Date Class Name Date Class MATH SKILLS TRANSPARENCY WORKSHEET 1 MATH SKILLS TRANSPARENCY
 WORKSHEET 2 Interpreting and Drawing Graphs Use with Chapter 2, Visualizing the Use with Chapter 3, Section 2.4 Conservation of Mass Section 3.2
 1. MATH SKILLS TRANSPARENCY MASTER Visualizing the ... Pages ... Name Date Class Calculating Numbers of Electrons Use with Chapter 7, Section 7.1
 and Predicting Ionic Change 1. What happens to a neutral atom if it loses one electron? snow.dearbornschools.org Math Skills Transparency 5. Math
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 master 5 interpreting waves use. Math Skills Transparency 5 Worksheets - Kiddy Math 6. Draw a circle graph to represent your answer to Question 7. 7.
 Complete the following statement by underlining the correct words in parentheses. To calculate percent by mass, you first divide a part (a whole) by
 (a part. he who 67 Then you multiply by Math Skills Transparency Worksheets Chemistry: Matter and Change Chapter
 3 oakman.dearbornschools.org Name Date Class 14 Use with Chapter 10, Section 10.3 MATH SKILLS TRANSPARENCY WORKSHEET Calculating the
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 question 3. Do your calculations support your answer in question 2? 6. If wave A has a frequency of 4.60 $\times 10^{14}$ s⁻¹, what is its wavelength in
 nanometers? Show your work. Interpreting Waves MATH SKILLS TRANSPARENCY WORKSHEET Use with Chapter 5, Section 5.1 5 MATH SKILLS
 TRANSPARENCY MASTER 5 Interpreting Waves Use ... MATH SKILLS TRANSPARENCY WORKSHEET Use with Chapter 25, Section 25.3 42 82 1. In
 equation 1, what do the numbers 238 and 92, written to the left of the symbol U, represent? 238 is the mass number of the isotope, and 92 is the
 atomic number of the isotope. 2. Explain how both mass number and atomic number have been conserved in equation 1. Name Date Class MATH
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 Matter and Change • Chapter 4 95 1. Look at the data table. What do the numbers 6 and 7 in 6X and 7X represent? 2. Look at step 1. What does amu
 stand for? What does it mean? 3. Look at step 2. Why is each isotope's mass multiplied by the isotope's percent abundance? 4. Name Date Class
 MATH SKILLS TRANSPARENCY MASTER 4 ... MATHSKILLSTRANSARENCY WORKSHEET Using the Periodic Table 1. Identify the number of valence
 electrons in each of the following elements. a. Ne d. Mg f. Cl h. Si 2. Identify the energy level of the valence electrons in each of the following
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 three worksheets that we had to do over the weekend! Just in case you missed out on some of the questions, here are the answers : Worksheet 1 :
 Determining Electronegativity difference and Percent Ionic Character.

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MATHSKILLSTRANSARENCY WORKSHEET Using the Periodic Table 1. Identify the number of valence electrons in each of the following elements. a. Ne d. Mg f. Cl h. Si 2. Identify the energy level of the valence electrons in each of the following elements. a. c. d. e. H Ar 1 3.

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6. Draw a circle graph to represent your answer to Question 7. 7. Complete the following statement by underlining the correct words in parentheses.

To calculate percent by mass, you first divide a part (a whole) by (a part) then you multiply by Math Skills Transparency Worksheets

Chemistry: Matter and Change Chapter 3

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MATH SKILLS TRANSPARENCY WORKSHEET Use with Chapter 25, Section 25.3 42 82 1. In equation 1, what do the numbers 238 and 92, written to the left of the symbol U, represent? 238 is the mass number of the isotope, and 92 is the atomic number of the isotope. 2. Explain how both mass number and atomic number have been conserved in equation 1.

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Name Date Class Calculating Numbers of Electrons Use with Chapter 7, Section 7.1 and Predicting Ionic Change 1. What happens to a neutral atom if it loses one electron?

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